



Year 12: Chemistry (Strand 1)						
	Autumn 1	Autumn 2	Spring 1	Spring 2	Summer 1	Summer 2
	Atomic Structure and Moles	Acids and REDOX	Periodicity	Enthalpy Changes	Rates of Reaction	Revision and Year 13 work.
Content: What will students know	<p>This section builds directly from GCSE Science, starting with basic atomic structure and isotopes. Important basic chemical skills are developed: writing chemical formulae, constructing equations and calculating chemical quantities using the concept of amount of substance.</p> <p>The role of acids, bases and salts in chemistry is developed in the context of neutralisation reactions.</p> <p>Finally, redox reactions are studied within the context of oxidation</p>	<p>The role of acids, bases and salts in chemistry is developed in the context of neutralisation reactions.</p> <p>Finally, redox reactions are studied within the context of oxidation number and electron transfer.</p>	<p>Periodic trends are first studied to extend the understanding of structure and bonding. Group properties are then studied using Group 2 and the halogens as typical metal and non-metal groups respectively, allowing an understanding of redox reactions to be developed further.</p> <p>Finally, this section looks at how unknown ionic compounds can be analysed and identified using</p>	<p>This section introduces physical chemistry within the general theme of energy. Learners first learn about the importance of enthalpy changes, their uses and determination from experimental results including enthalpy cycles.</p> <p>This section then investigates the ways in which a change in conditions can affect the rate of a chemical reaction, in terms of activation</p>	<p>The integrated roles of enthalpy changes, rates, catalysts and equilibria are considered as a way of increasing yield and reducing energy demand, improving the sustainability of industrial processes.</p>	<p>Students will spend some of this half term revising for mock examinations and preparing for Year 13 by further study of rates of reactions, orders, the rate equation and initial/graphical methods of studying reaction rates.</p>



Year 12: Chemistry (Strand 2)

	Autumn 1	Autumn 2	Spring 1	Spring 2	Summer 1	Summer 2
	Electron structure Bonding	Organic Chemistry introduction	Structure and reactions of alkanes and alkenes	Structure and reactions of alcohols and haloalkanes.	Instrumental Analysis	Revision and Year 13 Work
Content: What will students know	<p>This section introduces the concept of atomic orbitals and develops a deeper understanding of electron configurations linked to the periodic table. The central role of electrons in ionic and covalent bonding is then studied. The important role of molecules is studied, including an explanation of polarity and intermolecular forces. Finally, this section looks at how bonding and structure contribute to properties of substances.</p>	<p>This section is fundamental to the study of organic chemistry. This section introduces the various types of structures used routinely in organic chemistry, nomenclature, and the important concepts of homologous series, functional groups, isomerism and reaction mechanisms using curly arrows.</p>	<p>The initial ideas are then developed within the context of the hydrocarbons: alkanes and alkenes including free radical substitution and electrophilic addition reactions.</p>	<p>This section introduces two further functional groups: alcohols and haloalkanes, and considers the importance of polarity and bond enthalpy to organic reactions. Throughout this section, there are many opportunities for developing organic practical skills, including preparation and purification of organic liquids.</p>	<p>Finally, the important techniques of infrared spectroscopy and mass spectrometry are used to illustrate instrumental analysis as a valuable tool for identifying organic compounds.</p>	<p>Students will be preparing for examinations and then proceed to study Year 13 material on Benzene and its derivatives.</p>

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Assessment	End of topic assessment.	End of topic assessment.	End of topic assessment and mocks.	End of topic assessment.	End of topic assessment.	End of year assessment.
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Year 13: Chemistry (Strand 1)						
	Autumn 1	Autumn 2	Spring 1	Spring 2	Summer 1	Summer 2
	Rates of reactions and equilibria.	Acid dissociation constants and Buffers	Energy and Entropy	Transition Metals	Revision	Students on Study Leave
Content: What will students know	The largely qualitative treatment of reaction rates and equilibria encountered in Module 3 is developed within a quantitative and graphical context.	This section also allows learners to develop practical quantitative techniques involved in the determination of reaction rates and pH. There are many opportunities for developing mathematical skills, including use of logarithms and exponents, when studying the content of this section and when carrying out quantitative practical work.	Born–Haber cycles are used as a theoretical model to illustrate the energy changes associated with ionic bonding. Entropy and free energy are then introduced as concepts used to predict quantitatively the feasibility of chemical change. Redox chemistry permeates chemistry and the introductory work in Module 2 is developed further within this section, including use of	This section provides learners with a deeper knowledge and understanding of the periodic table within the context of the transition elements. This section includes the role of ligands in complex ions, stereochemistry, precipitation, ligand substitution and redox reactions. The colour changes and observations in these reactions increase the toolkit of qualitative inorganic tests for identifying unknown	Students will spend most of this time revising for their external examinations.	

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			volumetric analysis for redox titrations and an introduction of electrochemistry in the context of electrode potentials.	ionic compounds.		
Assessment	End of topic assessment.	End of topic assessment.	End of topic assessment and mocks	End of topic assessment.	End of topic assessment	A level Exams

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Year 13: Chemistry (Strand 2)						
	Autumn 1	Autumn 2	Spring 1	Spring 2	Summer 1	Summer 2
	Benzene and its derivatives	Carbonyl Compounds	Nitrogen compounds	Analysis and Combined Techniques	Revision	Students on Study Leave
Content: What will students know	This section extends the range of functional groups encountered in Module 4. Aromatic compounds are first introduced, including the central role of delocalisation within the chemistry of arenes and phenols. Directing groups are also introduced, including their importance to organic synthesis.	The important carbonyl compounds, aldehydes and ketones are then studied. Finally, carboxylic acids and their related functional groups, acyl chlorides and esters, are studied. The importance of acyl chlorides in organic synthesis is emphasised.	This section focuses on organic nitrogen compounds, including amines, amides and amino acids. Chirality and optical isomerism is also introduced. Condensation polymerisation is also introduced and compared with additional polymerisation. The importance of carbon-carbon bond formation in organic synthesis is stressed. Learners are also able to consider multi-stage	This section develops and complements the spectroscopic areas of organic chemistry previously encountered. This section demonstrates how analytical techniques introduced in Module 4 (infrared spectroscopy, mass spectrometry and elemental analysis) may be used in combination with NMR spectroscopy to provide evidence of structural	Students will spend most of this time revising for their external examinations.	

